

Calculations for Water Main Disinfection:

Given a 500 ft section of 8 inches water main, how much 65% Available Chlorine (HTH) is needed to provide an initial application of 50 mg/L?

(HTH = High Test Hypochlorite = Calcium Hypochlorite)

1. Convert 8" to feet (ft) $= (8 \text{ in} / 12 \text{ in} = 0.667 \text{ ft})$.
Diameter (d) = 0.667 ft.
Radius (r) = 0.333 ft ($\frac{1}{2}$ of diameter)
2. Find the volume (V) in cubic feet (ft³)
Volume = area of cross section of pipe (A) x Length (L)
 $V = A \times L$
Area = $(\pi) \times \text{the radius squared} (r^2 = r \text{ times itself})$
 $A = \pi \times r^2$
It is given that $\pi = 3.14$
 $V (\text{ft}^3) = 3.14 \times (r^2) \times \text{Length} \text{ or } 3.14 \times (0.333 \times 0.333) \times 500 = \underline{174.1 \text{ ft}^3}$
3. Find the amount of gallons of water in the line.
Conversion factor for water equals 7.48 gal per cubic feet (ft³)
Gallons equal Volume (ft³) x Conversion Factor for Water [Gallons/ (ft³)]
Gallons = $174.1 \text{ ft}^3 \times 7.48 \text{ gallons/ (ft}^3)$
Gallons = 1,302.2 gallons
4. Find lbs of 100% Available Chlorine
Available Chlorine is the actual chlorine available for disinfection
lbs of 100% Available Chlorine = Volume x Dose x Density of Water
Volume = Millions of Gallons (MG) = $1302.2 / 1,000,000 = 0.0013 \text{ MG}$
Dose = the concentration of the disinfectant in the water (mg/L) = 50 mg/L
Density of Water = 8.34 lbs/gallon
lbs of 100% Available Chlorine = $0.0013 \times 50 \times 8.34 = \underline{0.54 \text{ lbs}}$
5. Find lbs of 65% Available Chlorine.
lbs of 65% Available Chlorine =
lbs of 100% Available Chlorine/0.65 Available Chlorine
lbs of 65% Available Chlorine = $0.54 \text{ lbs} / 0.65 \text{ Available Chlorine} =$
0.83 lbs of 65 % Available Chlorine
6. Find ounces of 65% Available Chlorine.
ounces of 65% Available Chlorine =
lbs of 65% Available Chlorine x 16 oz/lb
ounces of 65% Available Chlorine = $0.83 \text{ lbs} \times 16 \text{ oz/ lb} = \underline{13.3 \text{ oz}}$

The following information is for water flowing at a velocity of 2.5 ft/sec through the indicated water main diameter:

MAIN SIZE	GALLONS PER MINUTE
2"	24
4"	96
6"	220
8"	392
10"	612
12"	880

Example:

Given a 500 ft section of 8 in water main, what volume of water is equal to a flow velocity of 2.5 ft/sec?

- Convert 8" to feet = (8 in / 12 in = 0.667 ft).
Diameter (d) = 0.667 ft.
Radius (r) = 0.333 ft (1/2 of diameter)
- Find: Volumetric Flow as cubic feet per second (ft³/sec)
Volumetric Flow (ft³/sec) =
Velocity (ft/sec) x Cross Sectional Area (ft²)
Velocity = 2.5 ft/sec
Given: $\pi = 3.14$
Cross Sectional Area =
 $\pi \times r^2 = 3.14 \times 0.333 \times 0.333 = 0.348 \text{ ft}^2$
Volumetric Flow = $2.5 \times 0.348 = \underline{0.87 \text{ ft}^3/\text{sec}}$
- Find Volumetric Flow as gal/sec:
Conversion factor for water equals =
7.48 gal per cubic foot (ft³)
gal/sec = ft³/sec x 7.48 gal/ft³ =
 $0.87 \times 7.48 = \underline{6.5 \text{ gal/sec}}$
- Find Volumetric Flow as gal/ min
gal/min = gal/sec x 60 sec/min =
 $6.5 \times 60 = \underline{390 \text{ gal/min}}$

Disinfection With Clorox

This method utilizes a "relational factoring" type of equation to determine the volume of Clorox (b) needed to give an initial concentration of disinfectant (C) in a particular volume (V) of water. The equation is based on mixing **1 fluid ounce of Clorox in 1 gallon of water** which yields an initial concentration of 410 mg/L per gallon or [(410 mg/L) (gallon)]. Clorox has a sodium hypochlorite (NaOCl) strength of about 5.25%.

If you know the volume of water to be treated and the desired initial concentration, then you may easily calculate the fluid ounces of Clorox required.

Example 1:

How many fluid ounces of Clorox must be added to 10,000 gallons of water to achieve an initial concentration of 410 mg/L?

$$\begin{aligned} b &= 10,000 \text{ gallons} \times 410 \text{ mg/L} \times \frac{\text{fluid ounces}}{(410 \text{ mg/L})(\text{gallons})} \\ b &= \underline{10,000 \text{ fluid ounces}} \end{aligned}$$

$$\text{OR } 10,000 \text{ fluid ounces} / 128 \text{ fluid ounces per gallon} = \underline{78 \text{ gallons}}$$

Example 2:

How many fluid ounces of Clorox must be added to 1,000 gallons of water to achieve an initial concentration of 50 mg/L?

$$\begin{aligned} b &= 1,000 \text{ gallons} \times 50 \text{ mg/L} \times \frac{\text{fluid ounces}}{(410 \text{ mg/L})(\text{gallons})} \\ b &= \underline{122 \text{ fluid ounces}} \end{aligned}$$

$$\text{OR } 122 \text{ fluid ounces} / 32 \text{ fluid ounces per quart} = \underline{3.8 \text{ quarts}}$$

It's as simple as that!